

**COMPOSITE STRUCTURE FOR HIGH EFFICIENCY HYDROGEN
SEPARATION AND ITS ASSOCIATED METHODS
OF MANUFACTURE AND USE**

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REFERENCE TO PROVISIONAL PATENT APPLICATION

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The matter of this application corresponds to the matter contained in Provisional Patent Application No. 60/475,620 entitled Composite Nano-Structure For High Efficiency Hydrogen Separation, which was executed on May 29, 2003, wherein this application assumes priority of that document.

BACKGROUND OF THE INVENTION

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1. FIELD OF THE INVENTION

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The present invention relates to systems and methods that are used to separate molecular hydrogen from a volume of gas. More particularly, the present invention is related to systems and methods that separate hydrogen from a volume of mixed gas by exposing the mixed gas to a hydrogen permeable material through which only atomic

hydrogen can readily pass.

2. DESCRIPTION OF THE RELATED ART

In industry there are many applications for
5 the use of molecular hydrogen. However, in many
common processes that produce hydrogen, the
hydrogen gas produced is not pure. Rather, when
hydrogen is produced, the resultant gas is often
contaminated with water vapor, hydrocarbons and/or
10 other contaminants. In many instances, however, it
is desired to have ultra pure hydrogen. In the art,
ultra pure hydrogen is commonly considered to be
hydrogen having purity levels of at least 99.999%.
In order to achieve such purity levels, hydrogen
15 gas must be actively separated from its
contaminants.

In the prior art, one of the most common ways
to purify contaminated hydrogen gas is to pass the
gas through a conduit made of a hydrogen permeable
material, such as palladium or a palladium alloy.
20 As the contaminated hydrogen gas passes through the

conduit, atomic hydrogen permeates through the walls of the conduit, thereby separating from the contaminants. In such prior art processes, the conduit is kept internally pressurized and is

5 typically heated to at least three hundred degrees centigrade. Within the conduit, molecular hydrogen disassociates into atomic hydrogen on the surface of the conduit and the conduit absorbs the atomic hydrogen. The atomic hydrogen permeates through the

10 conduit from a high pressure side of the conduit to a low pressure side of the conduit. Once at the low pressure side of the conduit, the atomic hydrogen recombines to form molecular hydrogen. The molecular hydrogen that passes through the walls of

15 the conduit can then be collected for use. One such system is shown by U.S. Patent No. 5,614,001 to Kosaka et al., entitled Hydrogen Separator, Hydrogen Separating Apparatus And Method For Manufacturing Hydrogen Separator.

20 The flow rate of hydrogen gas through the walls of a conduit is proportional to the length of

the conduit and the thickness of the walls of the conduit. Thus, a highly efficient purification system would have a very long, very thin conduit to maximize flow rate. However, palladium is a very
5 expensive precious metal. Consequently, conduits made of palladium and palladium alloys are very expensive to manufacture. As such, it is desirable to use as little of the palladium as possible in manufacturing a hydrogen gas purification system.
10 Furthermore, conduits made from palladium and palladium alloys typically hold gas under pressure and at high temperatures. Accordingly, the walls of the conduit cannot be made too thin, else the conduit will either rupture or collapse depending
15 on the pressure gradient present across the wall of the conduit.

A typical prior art conduit made from palladium or a palladium alloy would have a wall thickness of approximately $80\mu\text{m}$. The thickness of
20 the wall of the conduit is inversely proportional to the amount of purified hydrogen that passes

through that wall in a given period of time. As such, in order to make the conduit more efficient, a thinner wall is desirable. However, as has already been stated, a conduit wall cannot be made
5 so thin that it ruptures or collapses under the pressure of the gasses being passed through that conduit.

To further complicate matters, conduits made from palladium and palladium alloys may become less efficient over time as the interior walls of the conduits become clogged with contaminants. In order to elongate the life of such conduits, many manufacturers attempt to clean the conduits by reverse pressurizing the conduits. In such a
10 procedure, the exterior of the conduit is exposed to pressurized hydrogen. The hydrogen passes through the conduit wall and into the interior of the conduit. As the hydrogen passes into the interior of the conduit, the hydrogen may remove
15 some of the contaminants that were deposited on the interior wall of the conduit.
20

Due to the generally cylindrical shape of most prior art hydrogen purification conduits, the conduits are capable of withstanding a fairly high pressure gradient when the interior of the conduit
5 is pressurized higher than the exterior of the conduit. However, when such conduits are cleaned and the external pressure of the conduit is raised higher than the interior pressure, a much lower pressure gradient must be used, else the conduit
10 will implode.

In the prior art, improved conduit designs have been developed that attempt to minimize the amount of palladium used in a conduit, yet increase the strength of the conduit. One such prior art device is shown in U.S. Patent No. 4,699,637 to Iniotakis, entitled Hydrogen Permeation Membrane.
15 In the Iniotakis patent, a thin layer of palladium is reinforced between two layers of mesh. The laminar structure is then rolled into a conduit.
Such a structure uses less palladium. However, the
20 conduit is incapable of holding the same pressure

gradient as solid palladium conduits. Accordingly,
the increase in efficiency provided by the thinner
palladium layer is partially offset by the
decreased pressure limits, and thus gas flow rate,
5 that are capable of being processed.

Another prior art approach to limiting the
amount of palladium used is to create membranes
that are placed over apertures, like a skin on a
drum. A pressure gradient is then created on
10 opposite sides of the membrane, thereby causing
hydrogen to flow through the membrane. One such
system is shown in U.S. Patent No. 5,734,092 to
Wang et al., entitled Planar Palladium Structure. A
problem associated with such prior art systems is
15 that the palladium or palladium alloy membrane is
typically positioned in a level plane, wherein a
pressure gradient exists from one side of the
membrane to the other. Since the membrane is flat,
it has little structural integrity when trying to
20 resist the forces created by the pressure gradient.
Accordingly, in order to prevent the membrane from

rupturing, solid perforated substrates are used to reinforce the membrane. The solid perforated substrates, however, are complicated to manufacture, restrict the flow through the
5 membrane, and reduce the efficiency of the overall system.

U.S. Patent No. 6,152,987 to Ma, entitled Hydrogen Gas-Extraction Module And Method Of Fabrication, discloses a hydrogen separator where a
10 solid layer of hydrogen permeable material is deposited over a porous substrate of dissimilar material. The porous substrate supports the hydrogen permeable material and provides much more support than prior art mesh support systems.
15 However, the porous substrate only allows gas to contact the hydrogen permeable material where a pore gap is exposed to the hydrogen permeable material. This configuration greatly limits the area of hydrogen permeable material actually exposed to gas. Furthermore, due to differences
20 thermal coefficients and other physical properties,

hydrogen permeable material deposited on a
substrate of a dissimilar material tends to
separate from the substrate. This can cause leakage
of contaminated gas through the hydrogen permeable
material and the eventual failure of system.

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U.S. Patent Application No. 2003/0190486 to
Roa et al, also discloses a hydrogen separator
where a solid layer of hydrogen permeable material
is deposited over a porous substrate of dissimilar
material. In the Roa application, a palladium alloy
is deposited over the porous substrate of
dissimilar material using first electroplating
process. Copper is then deposited on the palladium
in a second electroplating process. The palladium
and copper layers are then annealed to produce a
palladium alloy in place on the substrate. However,
the porous substrate only allows gas to contact the
hydrogen permeable material where a pore gap is
exposed to the hydrogen permeable material. This
configuration greatly limits the area of hydrogen
permeable material actually exposed to gas.

Furthermore, palladium alloy does not bond well to the porous substrate of dissimilar material. This leads to eventual failure.

In prior art systems such as the Ma patent and
5 the Roa application, where palladium alloy is deposited directly on a porous substrate, another problem is that the substrates have large pore sizes to maximize exposure of the palladium alloy to gas. This requires that thick uneven layers of
10 palladium be deposited over the porous substrate to cover the pores of the substrate. This causes faults in the layer of palladium that tend to fail over time.

A need therefore exists in the art of hydrogen
15 purification for a system and method that can handle high flow rates of gas, per unit area, and yet uses only a minimal amount of hydrogen permeable material. A need also exists for a hydrogen purification system with a hydrogen
20 permeable layer that is strongly bonded to its supporting substrate so that it can more reliably

withstand large pressure gradients and repeated temperature cycles over longer periods of time.

These needs are met by the present invention as described and claimed below.

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SUMMARY OF THE INVENTION

The present invention is a hydrogen purification system and method that utilizes a novel composite hydrogen separator. The composite hydrogen separator has a first porous layer of a hydrogen permeable material, such as a palladium alloy. A solid layer of the same hydrogen permeable material is then disposed onto the first porous

15 layer. A pressure differential is created across the structure of the composite hydrogen separator.

The porous layer of hydrogen permeable material supports the solid layer and enables the solid layer to withstand large pressure differentials.

20 Furthermore, the porous layer of the hydrogen permeable material is diffusion bonded to the solid

layer, thereby greatly increasing the effective surface area of the solid layer that is exposed to hydrogen gas. Accordingly, a large flow rate of hydrogen gas can be obtained in a small amount of space.

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The hydrogen permeable material can be further supported by a porous metallic structure that adds structural integrity to the composite hydrogen separator and prevents the composite hydrogen separator from deforming when exposed to very large pressure gradients.

10

The result is a hydrogen purification system that is small and uses limited amounts of precious metals, yet has a high hydrogen gas flow rate.

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BRIEF DESCRIPTION OF THE DRAWINGS

For a better understanding of the present invention, reference is made to the following description of exemplary embodiments thereof, considered in conjunction with the accompanying drawings, in which:

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FIG. 1 is a schematic of an exemplary embodiment of a hydrogen purification system;

5 FIG. 2 is a fragmented view of an exemplary embodiment of a composite hydrogen separator with a cross-sectional enlargement;

10 FIG. 3 is a cross-sectional view of the base layer used in an exemplary embodiment of the present invention;

15 FIG. 4 is a schematic of a manufacturing step used in production of an exemplary embodiment of the present invention;

FIG. 5 is a schematic of a manufacturing step used in production of an exemplary embodiment of the present invention;

20 FIG. 6 is a schematic of a manufacturing step

used in production of an exemplary embodiment of
the present invention;

5 FIG. 7 is a schematic of a manufacturing step
used in production of an exemplary embodiment of
the present invention;

10 FIG. 8 is a schematic of a manufacturing step
used in production of an exemplary embodiment of
the present invention;

15 FIG. 9 is a schematic of an alternate
manufacturing process used in production of an
alternate exemplary embodiment of the present
invention;

20 FIG. 10 is a cross-sectional view of an
alternate embodiment of a composite hydrogen
separator;

FIG. 11 is a schematic of a second exemplary

manufacturing process used to form the composite
hydrogen separator of the present invention; and

5 FIG. 12 is a schematic of a third exemplary
manufacturing process used to form the composite
hydrogen separator of the present invention.

DETAILED DESCRIPTION OF THE ILLUSTRATED EMBODIMENTS

10 The present invention system provides a means
for purifying a hydrogen gas at a high flow rate,
using a small amount of space and a small amount of
precious metals.

15 Referring to Fig. 1, there is shown a
schematic of an exemplary embodiment of a hydrogen
purification system 10 in accordance with the
present invention. The hydrogen purification system
10 has a collection chamber 12 for collecting
purified hydrogen gas. At least one supply conduit
20 14 extends into the collection chamber 12 where
each of the supply conduits 14 is coupled to a

source of contaminated hydrogen gas 16. In one embodiment of the present invention, the supply conduit 14 is fabricated from stainless steel or an equivalent alloy that is capable of retaining the contaminated hydrogen gas at a first pressure and at a predetermined operating temperature of at least three hundred degrees centigrade.

Along the supply conduit 14 is disposed a composite hydrogen separator 20. The composite hydrogen separator 20 enables only pure hydrogen gas to permeate out of the supply conduit 14 and into the collection chamber 12. Once, the purified hydrogen enters the collection chamber 12 it is collected for use.

The composite hydrogen separator 20 is the only part of the purification system 10 that is permeable to hydrogen gas. As such, the flow capacity of the overall purification system 10 relies completely upon the flow capacity characteristics of the composite hydrogen separator 20.

Referring to Fig. 2, it can be seen that the example of the composite hydrogen separator 20 that is illustrated is tubular in shape, wherein it defines an open central conduit 22. The composite 5 hydrogen separator 20 is attached to the supply conduit 14. Contaminated hydrogen gas under pressure is passed through the central conduit 22 in the center of the composite hydrogen separator 20. The pressure in the surrounding collection 10 chamber 12 is kept lower than the pressure within the central conduit 22. As such, a pressure differential exists across the wall 24 of the composite hydrogen separator 20 that causes hydrogen gas to permeate through the structure of 15 the composite hydrogen separator 20. The greater the pressure differential between the conduit 22 in the composite hydrogen separator 20 and the collection chamber 12, the greater the rate of flow 20 of hydrogen out of the composite hydrogen separator 20.

The direction of the flow can be controlled by

selectively changing the direction of the pressure differential between the collection chamber 12 and the conduit 22 within the composite hydrogen separator 20. Under normal operating conditions,
5 the pressure within the composite hydrogen separator 20 is kept greater than the pressure within the collection chamber 12. As such, hydrogen gas flows out of the composite hydrogen separator 20 and is purified. However, during maintenance cycles,
10 the pressure of hydrogen gas within the collection chamber 12 can be made greater than the pressure of hydrogen within the composite hydrogen separator 20. In this circumstance, the hydrogen gas flows from the collection chamber 12 into the composite hydrogen separator 20 and can be used to
15 remove contaminants that may have collected on the interior surfaces of the wall 24 of the composite hydrogen separator 20.

The wall 24 of the composite hydrogen separator 20 has a composite structure that is made using a unique combination of formation processes

and techniques. From the enlarged segment of Fig. 2, it can be seen that the composition of the wall 24 of the composite hydrogen separator 20 changes across the thickness of the wall 24. As such, the 5 wall 24 is a composite structure of different materials.

The first layer of the wall is the supporting base layer 30. The base layer 30 is made of a non-reactive metal or metal alloy that in itself is not 10 hydrogen permeable. In a preferred embodiment, the material selected from the base layer is tungsten or molybdenum. However, the base layer 30 of the wall 24 is porous and allows hydrogen gas to freely pass through the pores of its structure. As will 15 later be explained in more detail, the base layer 30 is made using traditional powder metallurgy techniques, where a powder of the non-reactive metal or alloy is formed into the desired shape and sintered into a porous form. Preferably, the base 20 layer 30 is manufactured so that it has pore size of approximately $0.7\mu\text{m}$ prior to sintering. However,

larger and smaller pore particle sizes can be used.

To promote strong bonding to the porous base layer 30, bonding layers 31 are deposited onto the base layer 30. The bonding layers 31 are thin so

5 that they do not close the pores in the base layer 30. The bonding layers are preferably comprised of three separate layers. The first bonding layer 33 is chromium. The chromium layer is deposited directly onto the tungsten of the base layer 30.

10 The first bonding layer 33 of chromium bonds with the tungsten of the base layer 30, as will later be explained.

The second bonding layer 35 is a layer of tungsten. The tungsten is deposited over the first bonding layer 33 of chromium and bonds to the chromium.

15 The third bonding layer 37 is a layer of copper or silver, depending upon the palladium alloy to be used in the hydrogen separator 20. If a 20 palladium-silver alloy is used, silver is used as the third bonding layer 37. However, it is

preferred that a palladium-copper alloy be used. As such, copper is preferably used as the third bonding layer 37. The silver or copper acts as a wetting layer in that it covers the second bonding layer 35 of tungsten and provides a surface onto which a palladium alloy can directly diffusion bond.

A first porous layer 32 of hydrogen permeable material is then formed upon the bonding layers 31. The first porous hydrogen permeable layer 32 is preferably made from a palladium alloy, such as $Pd_{1-x}Ag_x$ or $Pd_{1-x}Cu_x$. However, other suitable hydrogen permeable alloys can also be used. Such alloys include, but are not limited to $PdRu$, $PdAu$, $PdNi$, and $PdFe$. The first porous layer 32 can be made from particles of a single size. However, in a preferred embodiment, the size of the palladium alloy particles in the first porous layer are graduated. The palladium alloy particles become smaller the further those particles are from the bonding layers 31 and the base layer 30. The first

porous hydrogen permeable layer 32 can be formed in different ways, as will be later described.

A solid layer 34 of palladium alloy or other hydrogen permeable material is deposited over the first porous hydrogen permeable layer 32 using one of the techniques later described. The hydrogen permeable material selected for the solid layer 34 is preferably the same material selected for the first porous hydrogen permeable layer 32. In this manner, the solid layer 34 and the first porous hydrogen permeable layer 32 share physical characteristics, such as thermal expansion. Furthermore, by using the same hydrogen permeable material, the solid layer 24 of hydrogen permeable material can be diffusion bonded to the first porous hydrogen permeable layer 32.

The deposited solid layer 34 of hydrogen permeable material is preferably between about 0.05 microns and about 0.20 microns thick. The deposited solid layer 34 diffusion bonds with the first porous hydrogen permeable layer 32. However, the

deposited solid layer 34 seals the pores of the
first porous hydrogen permeable layer 32. As such,
any hydrogen gas traveling through the wall 24 of
the composite hydrogen separator 20 must permeate
5 through the deposited solid layer 34 in order to
advance.

Above the deposited solid layer 34 of hydrogen
permeable material is a second porous hydrogen
permeable layer 36. The second porous hydrogen
10 permeable layer 36 is also made of the same
hydrogen permeable material as the solid layer 34.
However, the second porous hydrogen permeable layer
36 is manufactured to be porous. As such, the
deposited solid layer 34 of hydrogen permeable
15 material is interposed in between two porous layers
32, 36 of hydrogen permeable material.

The base layer 30 of non-reactive metal at the
bottom of the wall 24 provides the wall 24 of the
composite hydrogen separator 20 with physical
20 strength. This base layer 30 prevents the composite
hydrogen separator 20 from being easily inwardly

crushed or outwardly expanded. Accordingly, the composite hydrogen separator 20 can withstand large pressure differentials without inwardly or outwardly deforming.

5 The first porous hydrogen permeable layer 32 and the second porous hydrogen permeable layer 36 serve two purposes. First, these layers physically support the deposited solid layer 34 of hydrogen permeable material on either side of the solid layer 34. Accordingly, the solid layer 34 of hydrogen permeable material is fully supported at nearly all points along its surface. However, although most all points of the solid layer 34 are supported, the entire solid layer 34 remains active

10 in the permeation of hydrogen gas because the supporting porous layers 32, 36 are also completely permeable to hydrogen gas. The solid layer 34 of hydrogen permeable material can therefore withstand very large pressure differentials between the sides

15 of the solid layer 34 without rupturing.

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 The first porous hydrogen permeable layer 32

and the second porous hydrogen permeable layer 36 do not obstruct the surface area of the solid layer 34 that is exposed to hydrogen gas. In fact, these porous layers 32, 36 actually increase the effective surface area of the solid layer 34. The first porous hydrogen permeable layer 32 and the second porous hydrogen permeable layer 36 are preferable made of the same material as the deposited solid layer 34. The contact of the first porous hydrogen permeable layer 32 and the second porous hydrogen permeable layer 36 against the solid layer 34 therefore has the same effect as manufacturing the solid layer 34 with highly textured surfaces. This increases the effective surface area that is exposed to hydrogen gas trying to permeate through the solid layer 34.

Thus, in operation, contaminated gas enters the conduit 22 in the center of the composite hydrogen separator 20. The contaminated gas passes through the pores in the base layer 30 of the wall 24. The contaminated gas also passes through the

pores in the first porous hydrogen permeable layer
32 until it contacts the deposited solid layer 34.
The deposited solid layer 34 blocks all gasses
except hydrogen. The hydrogen permeates through the
solid layer 34 where the hydrogen gas directly
contacts the solid layer 34. However, the hydrogen
gas can also permeate into the areas of the first
porous hydrogen permeable layer 32 that are
contacting and supporting the deposited solid layer
34. As a result, a very large surface area is
available through which hydrogen gas can pass.

Once the hydrogen gas permeates across the
deposited solid layer 34, it enters the second
porous hydrogen permeable layer 36. The second
porous hydrogen permeable layer 36 provides very
little resistance to the further passage of the now
purified hydrogen gas into the surrounding
collection chamber 12.

When the composite hydrogen separator 20 is
being cleaned, pressure differentials are reversed.
The second porous hydrogen permeable layer 36 then

acts to increase the effective surface area of the deposited solid layer 34 as hydrogen gas flows into the composite hydrogen separator 20 toward the central conduit 22.

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EXAMPLE OF A MICROSTRUCTURE COMPOSITION FOR A
HYDROGEN SEPARATION MEMBRANE

As has been previously mentioned, various layers of the wall 24 of the composite hydrogen separator 20 can be made in different ways.

Referring to Fig. 3, it will be understood that the base layer 30 is preferably a sintered metal that is sintered into a porous form having a pore size of approximately 0.5 μm to 0.9 μm . An average pore size of about 0.7 μm microns being preferred and used by way of example. Because of the overall process flexibility and geometric scaling, there is a considerable variation in possible average pore diameter. The base layer 30 can be sintered in the desired form 41 or a sintered block can be machined

into the desired base form 41. In the shown example, the desired base form 41 is a tubular section.

Referring to Fig. 4, it can be seen that
5 individual bonding layers 33, 35, 37 are preferably deposited onto the base layer 30 using biased substrate sputtering, where the base layer 30 is the substrate. The base form 41 is placed in a vacuum deposition chamber, wherein each of the
10 individual bonding layers 33, 35, 37 are deposited separately. In a preferred embodiment for this specific pore size of about 0.7 micron, the first bonding layer 33 of chromium is deposited in a layer of about 0.03 microns. This initial sputtered
15 Chromium layer promotes adhesion. The second bonding layer 35 of tungsten is deposited in a layer of about 0.10 microns. The second bonding layer 35 of sputtered tungsten is for a diffusion barrier between the first bonding layer 33 and the
20 subsequent third bonding layer 37. The third bonding layer 37 of copper is deposited in a layer

of about 0.20 microns. ~~Each of the individual bonding layers 33, 35, 37 having a variation tolerance of at least 50% of the deposited thickness.~~ The third bonding layer 37 of copper is 5 for diffusion bonded to the porous layers that follow the sputtering process. The thickness of these three initial bonding layers 33, 35, 37 can be adjusted in accordance with the pore size of layer 30. For the case of an average pore size of 10 about 0.7 μm , the thickness of the three bonding layers 33, 35, 37 is preferably what is stated in the example.

With the preferred thickness of the individual bonding layers 33, 35, 37 stated, the combined bonding layers 31 will have a combined thickness of 15 0.33 microns. The base layer 30 preferably has a pore size of about 0.7 microns. As the individual bonding layers 33, 35, 37 are deposited, the pore size present on the exterior of the assembly becomes smaller. The deposition of thin film 20 material onto facing pore surfaces, obtained

through substrate biasing, is controlled in a manner consistent with the specifics to the vacuum system geometry, power distribution, and operating parameters during plasma sputtering. In the case of
5 the example, the 0.7 micron pores will be reduced in correspondence to the bonding layer thickness of 0.33 microns; the resulting new pore size will be about 0.36 microns, or approximately a 50% reduction in diameter. However, the pores do remain
10 open after the combined bonding layers 31 are applied.

From the above discussion, the flexibility in the process can be seen as a direct correspondence between the initial pore size of base layer 30 and
15 the total thickness of the deposited bonding layers 33, 35, 37. Specifically, the larger the initial pore size in the base layer 30, the thicker the layer of deposited bonding layers 33, 35, 37, since the pore size can be controlled by the thickness of
20 the bonding layers 33, 35, 37. The thickness of the bonding layers 33, 35, 37 may be approximately

proportional to the original thickness stated above
for the case of pore size in base layer 30 of 0.7
microns.

It will be understood that the material of
5 described for the bonding layers 33, 35, 37 are
exemplary and can be changed. For instance, the
first two bonding layers 33, 35 can be made from
material that is not hydrogen permeable. The third
bonding layer 37 can be made permeable to hydrogen
10 and alloy-forming with subsequent layers for
diffusion bonding

Referring to Fig. 5, it can be seen that the
first porous hydrogen permeable layer 32 is
deposited over the bonding layers 33, 35, 37. This
15 can be accomplished in a variety of ways. In the
shown manufacturing method, particles of palladium
and copper are suspended in a rheological vehicle,
such as Heraeus-Cermalloy V633 or V636. Such
rheological vehicles are thixotropic agents which
20 are commonly used in nitrogen fireable thick film
paste formulations. This suspension forms a thick

film paste 43. Using traditional thick film deposition techniques, such as spin coating, spraying, or screen printing, the thick film paste 43 is deposited over the bonding layers 31,
5 creating a wet four layered slug 45. The particles of palladium and copper in the thick film paste 43 are larger than the pores within the exterior of the bonding layers 31. In this manner, the particles of palladium and copper in the thick film
10 paste 43 will not clog the existing pores.

After the layer of thick film paste 43 is applied, the wet slug 45 is dried and fired (sintered) at elevated temperatures into a first thick film layer 47. . The drying step is typically performed at 150 °C for 10 minutes to drive off the thinning constituents of the paste. The sintering step is performed at a temperature that is determined by the physical properties of the constituent particles in the thick film paste 43 in order to result in an alloy. In this case, the sintering can be performed at a temperature of 600°C
15
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for several hours in an ambient containing several percent of hydrogen. . Within the first thick film layer 47, the sintering induces diffusion bonding between the palladium and the copper particles,
5 producing a palladium alloy. The sintering also induces diffusion bonding between the bonding layers 31, the base layer 30 with the bonding layers 31, and the bonding layers 31 with the first thick film layer 47. This first thick film layer 47
10 is then bonded cohesively (sintered) and adhesively to the base layer 30 through bonding layers 31.

After annealing, a five layered slug 49 is created having a porous base layer 30, three bonding layers 31 and the first thick film layer 47
15 that are all diffusion bonded together.

Referring to Fig. 6, it can be seen that subsequent thick film layers can be deposited over the first thick film layer 47, thereby growing the thickness of that layer. The five layered slug 49
20 is repeatedly subjected to thick film deposition and sintering. Each time a subsequent thick film

layer is deposited, the size of the particles used in the thick film paste are decreased, thereby also decreasing the average pore between the particles.

The subsequent layers of thick film are sintered

5 and the process is repeated until a first porous hydrogen permeable layer 32 is produced. Hence, the first porous hydrogen permeable layer 32 is comprised of a number of sintered particle sub-layers, where the pore size of each subsequent sub-

10 layer is decreased with respect to the prior layer, progressively decreasing until the final sub-layer has a pore size that reaches a designed value.

The designed value of the pore size of the final sub-layer is in accordance with the thickness

15 of the solid palladium layer that will be used. As such, the microstructure of permeable layer 32 as a whole possesses the mechanical strength required to withstand the range of differential pressures across a solid palladium layer while in operation.

That is, for small values of differential pressure,

20 a solid palladium layer can be relatively thin,

hence the final porous layer should possess a relatively small average pore in order for the deposition of a solid palladium layer to close the pores. Similarly, for operating with higher differential pressures, a thicker solid palladium layer may be required for mechanical strength, which therefore requires a thicker solid film and a larger pore size can be used. The design of the microstructure accounts for the practical range of operation of a solid palladium layer, while maximizing the permeance of the solid palladium layer by maintaining a minimum thickness of a solid palladium layer for the required mechanical strength. The design is therefore scalable in accordance with the pressure differential required for mechanical strength.

The sub-layers produce a first porous hydrogen permeable layer 32 has a final thickness T_1 , which is the sum of the annealed thicknesses of all the thick film sub-layers that have been deposited. The first porous hydrogen permeable layer 32 is

comprised of a single palladium alloy. However, the
particle size and the pore size within the first
porous hydrogen permeable layer 32 decreases as a
function of distance from the base layer 30. In the
5 outer most application of thick film, it is
preferred that the pore size after annealing be
between 0.03 microns and 0.05 microns. This pore
size produces a relatively smooth surface upon
which a solid layer of material can later be
10 deposited.

Referring to Fig. 7, it can be seen that a
solid layer 34 of palladium alloy is deposited over
the first porous hydrogen permeable layer 32. Many
deposition techniques can be used, such as vapor
15 deposition, sputter deposition, spraying, spin
coating, dipping or electroplating, or electroless
plating. The minimum thickness of the solid layer
34 of palladium alloy is such that the pores are
essentially closed after the deposition of the
20 solid layer 34. For bias sputtering, the thickness
of the solid layer 34 can be about one-half the

average pore size diameter of first porous hydrogen
permeable layer 32 under the appropriate sputtering
conditions. However, it is preferred that the
thickness of the solid layer 34 of palladium be at
5 least as thick as the pore size existing on the
exterior of the first porous hydrogen layer 32 in
order to compensate for the deviation in the actual
pore. The solid layer 34 of palladium alloy seals
the pores of the first porous hydrogen permeable
10 layer 32, thereby preventing any contaminant gas
from flowing through the wall structure of the
hydrogen separator.

Since the size of the particles used in the
first porous hydrogen permeable layer 32 are
15 smallest at its exterior, the surface on which the
solid layer 34 is deposited is very even and
smooth. The solid layer 34 of palladium covering
this smooth layer can therefore be applied in an
even layer with relatively small variations and
20 imperfections.

Referring to Fig. 8, it can be seen that after

the solid layer 34 of palladium alloy is deposited,
at least one thick film layer of palladium and
copper particles are deposited onto the solid layer
34 of palladium alloy. After annealing, this
5 produces a second porous hydrogen permeable layer
36 that covers the solid layer 34 of palladium
alloy. The particle size within the second porous
hydrogen permeable layer 36 can also be graduated
by depositing multiple thick film layers onto the
10 solid layer 34 of palladium alloy. However, the
size of the pores in the second porous hydrogen
permeable layer 36 that lay against the solid layer
34 of palladium alloy are preferably the same as
the pore size in the first porous hydrogen
15 permeable layer 32 immediately adjacent the solid
layer 34 of palladium alloy.

The contact of the first porous hydrogen
permeable layer 32 and the second porous hydrogen
permeable layer 36 with the solid layer 34 of
20 palladium alloy increases the effective surface
area of the solid layer 34 that is exposed to

hydrogen gas diffusion. The purpose of this
increase can be explained by considering the
following example. Suppose that the powder particle
size of the hydrogen permeable material in both the
5 first porous hydrogen permeable layer 32 and the
second porous hydrogen permeable layer 36 has an
average diameter d on the surfaces that contact the
solid layer 34 of palladium alloy. In a simple
cubic lattice of connected particles, the surface
10 area occupied by each powdered particle would be d^2 .
The pore area created by the contact between
particles in each layer is $d^2(1-\pi/4)$. Furthermore,
the exposed area of each hemisphere of a particle
is $d^2(\pi/2)$. The sum of these areas is $d^2(1+\pi/4)$,
15 which equals $1.79d^2$. Consequently, the net effect of
interposing the deposited solid layer 34 of
hydrogen permeable material between porous layers
32, 36 is to increase the effective surface area of
the deposited solid layer 34 by seventy-nine
20 percent (79%).

The solid layer 34 of hydrogen permeable

material therefore has its surface areas effectively increased by approximately 79% by having the solid layer 34 supported by porous layers 32, 36 of the same hydrogen permeable material.

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In Fig. 3 through Fig. 8, the porous layers 32, 36 of hydrogen permeable material were described as being deposited as thick film layers using traditional thick film deposition techniques.

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Such techniques are preferred if the first and/or second porous layers of hydrogen permeable material are to have graduated particle and pore sizes.

However, the first and/or second porous layers of hydrogen permeable material can be manufactured with one consistent particle size and pore size. In such a construction other manufacturing techniques are available.

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Referring to Fig. 9, such an exemplary method of manufacture is illustrated. In Fig. 9, the material 40, such as tungsten, that is not hydrogen permeable is provided in powdered form having a

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known powder particle size. The powdered metal is processed in a first powder metallurgy process 42, where the powder is compressed into a desired form and sintered. The result is a base layer slug 44
5 having a porous wall.

Bonding layers are then applied to the base layer slug 44 using a repeated sputter deposition process 53. The bonding layers are preferably chromium, tungsten and copper, as has previously been described. This creates a four layer slug 55.
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The four layer slug 55 is then sent into a second powder metallurgy process 46. In this second process, powdered hydrogen permeable material is compressed around the four layer slug 55 and is sintered. This creates a five layer slug 48.
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The five layer slug 48 is then taken into a deposition process 50, where a solid layer of hydrogen permeable material is deposited around the five layer slug 48, therein creating a six layer slug 52. The deposition process 50 can be any known process by which a thin solid layer of hydrogen
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permeable material can be deposited onto the five layer slug 48 so as to coat the five layer slug 48.

The six layer slug 52 is then subjected to a third powder metallurgy process 54 where another 5 layer of powdered hydrogen permeable material is compressed around the six layer slug 52 and sintered. The result is a complete seven layered composite hydrogen separator 59.

In the manufacturing technique depicted in Fig. 9, powdered material is compacted and sintered to form the porous layers of hydrogen permeable material. The powdered material need not be homogenous metal dust. Rather, it is possible to provide inert ceramic nano-particles and coat those 15 ceramic nano-particles with palladium alloy or some other hydrogen permeable material. The coated ceramic nano-particles can then be compacted and sintered as though they were homogenous metal powder. The advantages of using coated ceramic 20 nano-particles is that the particle size of the powder can be kept much more consistent than can

solid metal powder. Thus, a more consistent pore size can be created in the porous layers.

In the embodiments of the composite hydrogen separator thus illustrated and described, a single solid layer is used. The solid layer is surrounded by a layer of hydrogen permeable material that forms the exterior of the hydrogen separator. Such configurations are merely exemplary and it should be understood that layers within the hydrogen separator can be repeated or eliminated. A composite hydrogen separator with several layers is therefore possible with multiple deposited solid layers disposed between porous layers.

In Fig. 10, an alternate embodiment for the wall 60 of a composite hydrogen separator is shown. In this embodiment, the composite wall 60 of the hydrogen separator only has six layers. The six layers are a base layer 62, the three bonding layers 63, a first hydrogen permeable layer 64 and a deposited solid layer 66. The base layer 62, the bonding layers 63 and the first porous hydrogen

permeable layer 64 are the same as those previously described. However, the deposited solid layer 66 is now the outermost layer. To increase the surface area of the exterior of the deposited solid layer

5 66, the exterior of the deposited solid layer 66 is contoured. The contouring can be created by masking sections of the solid layer 66 during deposition or etching the solid layer 66 after deposition. The contouring of the exterior of the deposited solid

10 layer 66 serves to increase the exterior surface area of the solid layer 66, thereby serving the same function as the second hydrogen permeable layer used in previous embodiments. However, it can be seen that the embodiment of Fig. 10 uses less

15 palladium alloy than the previous embodiments, and is thus less expensive.

Referring to Fig. 11, yet another alternate embodiment for the wall 70 of the present invention composite hydrogen separator is shown. In this embodiment, the composite wall 70 of the hydrogen separator again has six layers. The six layers are

a base layer 72, the three bonding layers 73, a porous hydrogen permeable layer 74 and a deposited solid layer 76. The base layer 72 can be made using the sintered powder techniques previously described. The bonding layers 73 are then deposited on the base layer 72. On top of the bonding layers 73 is deposited a layer of hydrogen permeable material. The layer is initially deposited as a solid layer. The layer is then perforated using traditional masking and etching techniques, thereby creating the porous hydrogen permeable layer 74 that is shown. Since the porous hydrogen permeable layer 74 is created using masking and etching techniques, the size of the holes created in the porous hydrogen permeable layer 74 can be precisely controlled.

A solid layer 76 of hydrogen permeable material is then deposited over the porous layer 74. The top of the solid layer 76 is then partially etched to increase the surface area of the solid layer 76 exposed to the surrounding atmosphere. The

etching can be accomplished either through chemical etching or laser etching.

In the embodiments of the present invention, the hydrogen separator has been described in the form of a tubular conduit. It will be understood that such an embodiment is merely exemplary. The present invention hydrogen separator can be created in any shape, provided the hydrogen separator acts as a barrier between an area of contaminated hydrogen gas and an area of purified hydrogen gas.

Referring to Fig. 12, a construction of a hydrogen separator 80 is shown that is not tubular in shape and using no powdered metallurgical techniques in its construction. In the embodiment of Fig. 12, a base substrate 82 is provided. The base substrate is made of material that is not permeable to hydrogen and can be either a metal alloy or a semiconductor. The base substrate 82, however, is made porous by chemically etching or laser etching. Bonding layers 83 are deposited on the base substrate 82. A first layer of hydrogen

permeable material 84 is applied to the bonding layers 83 using tradition deposition techniques. The first layer of hydrogen permeable material 84 is then perforated by chemical etching or laser etching. The size of the perforations in the base substrate 82 and the size of the perforations in the first layer of hydrogen permeable material 84 are preferably made to be comparable to maintain a constant porosity throughout the two layers.

10 A second layer of hydrogen permeable material 86 is then deposited over first layer of hydrogen permeable material 84. The second layer of hydrogen permeable material 86 can be partially etched to increase the surface area on exposed on its top.

15 However, the etching of the second layer of hydrogen permeable material 86 is only partial and the second layer of hydrogen permeable material 86 remains as a solid barrier over the first porous layer of hydrogen permeable material 84.

20 It will be understood that the embodiments of the present invention system and method that are

described and illustrated herein are merely exemplary and a person skilled in the art can make many variations to the embodiment shown without departing from the scope of the present invention.

5 All such variations, modifications and alternate embodiments are intended to be included within the scope of the present invention as defined by the appended claims.